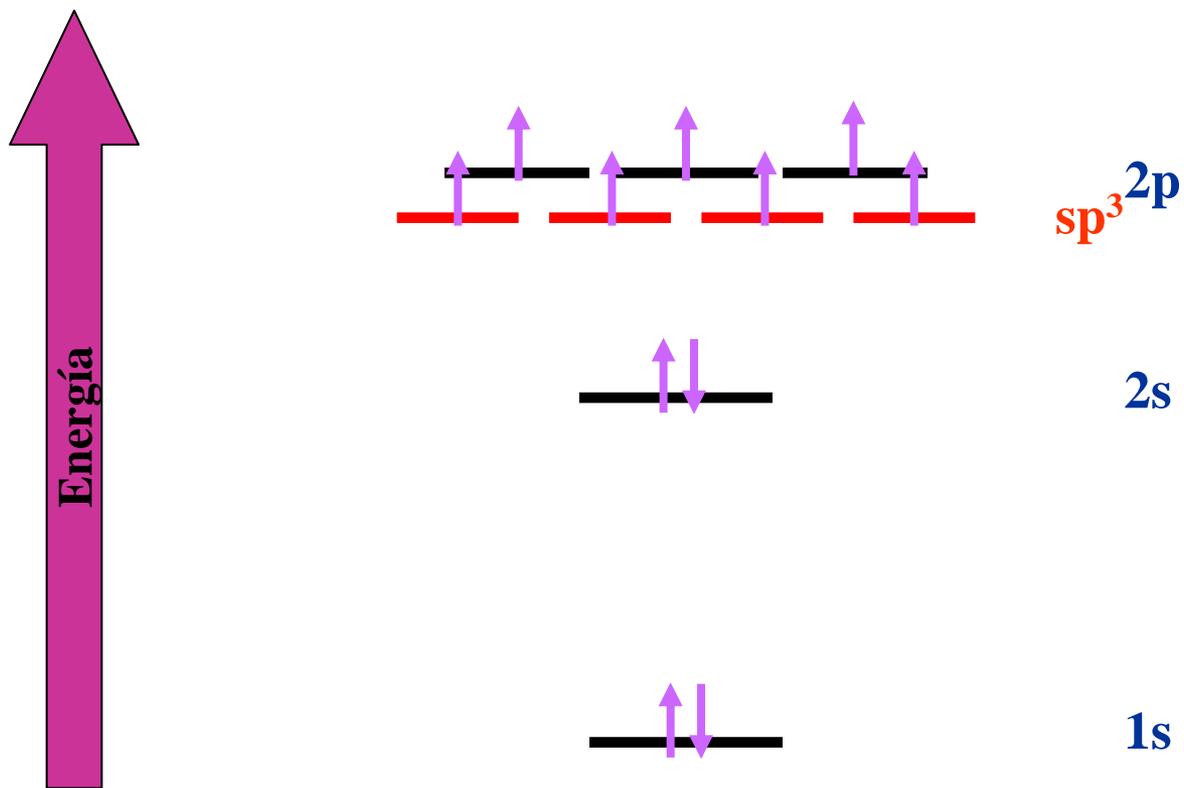
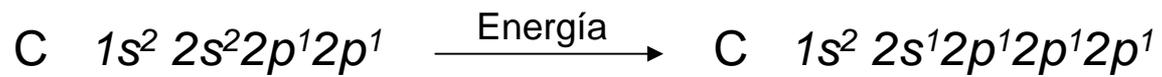
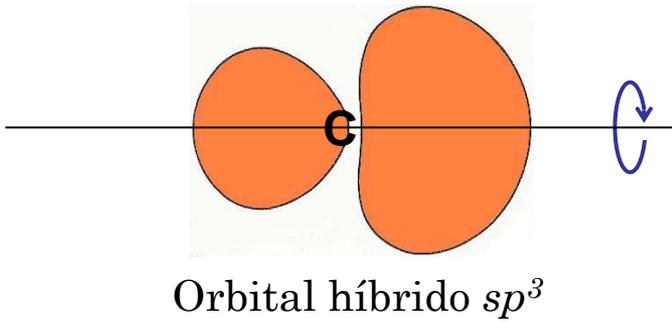


Hibridación de OA

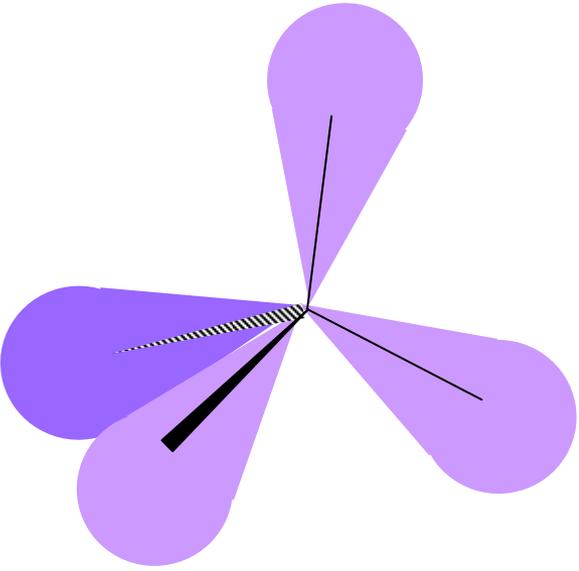
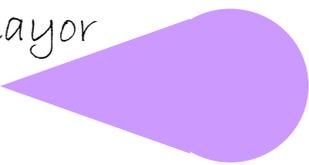
Hibridación sp^3 y las estructuras de metano y etano.





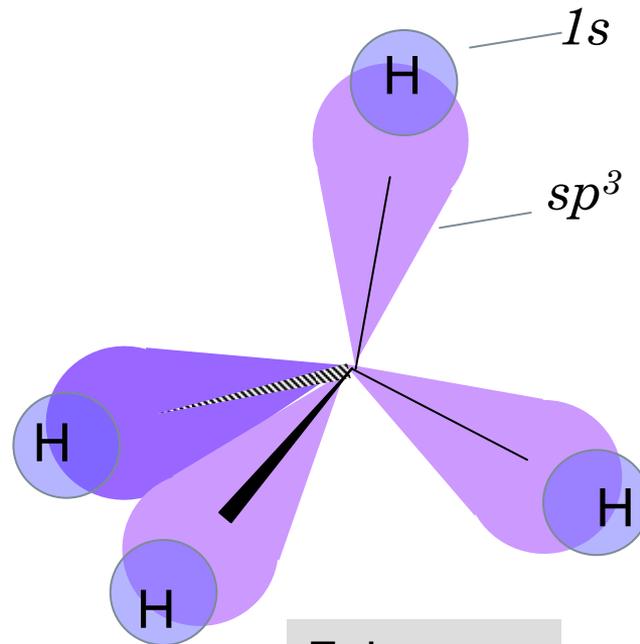
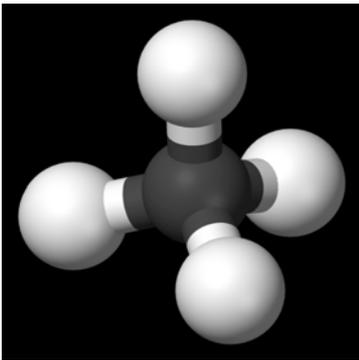
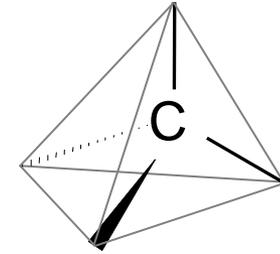
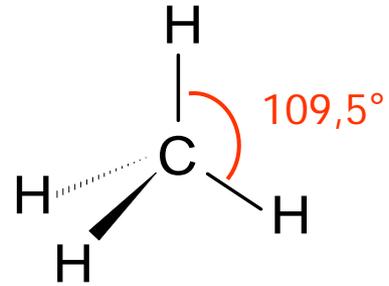
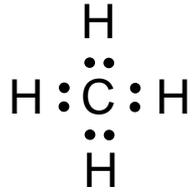
- 2 lóbulos
- Direccional
- Simetría de revolución
- Disposición tetraédrica de los cuatro ejes.

aunque para simplificar el dibujo, representaremos sólo el lóbulo mayor y de manera estilizada.

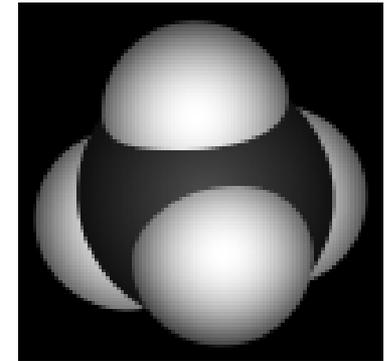


Disposición espacial de los 4 orbitales híbridos del C tetraédrico

Metano



Enlaces σ





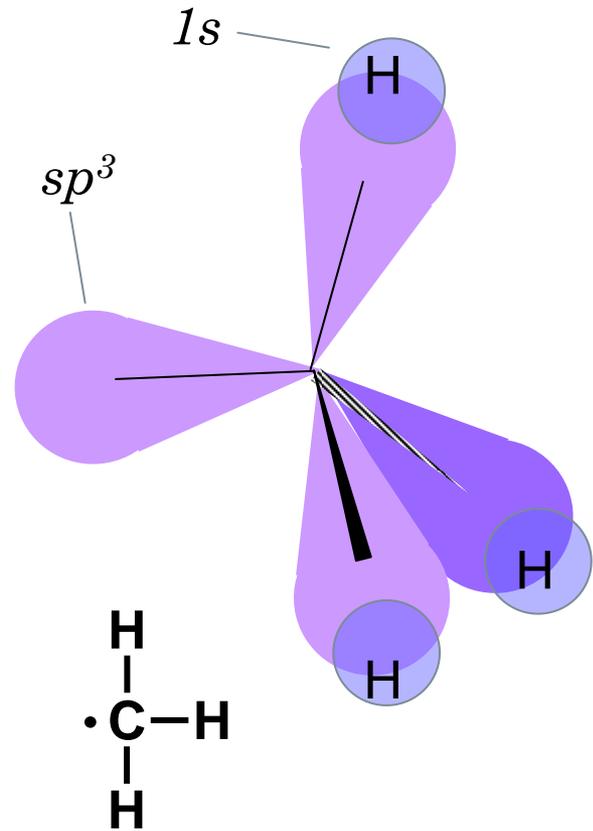
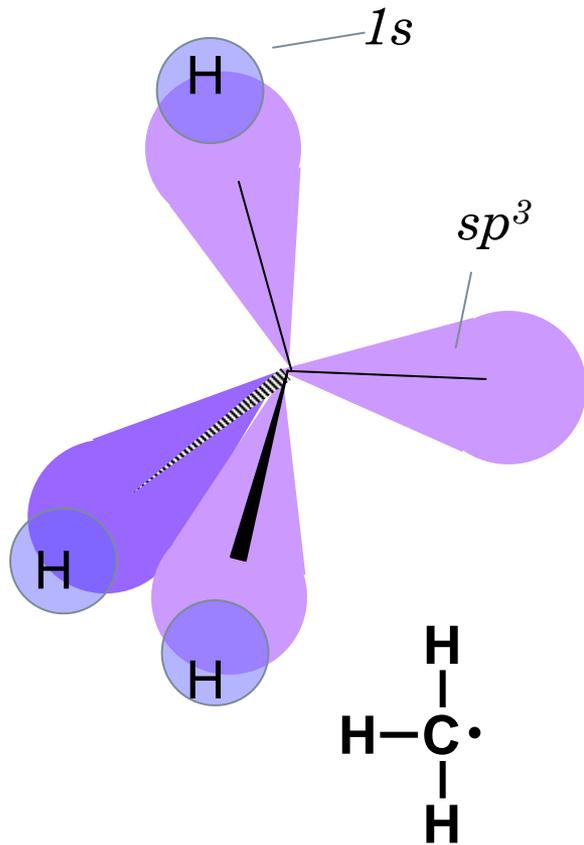
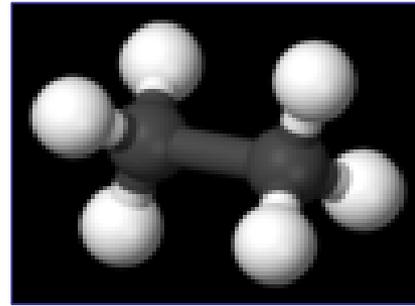
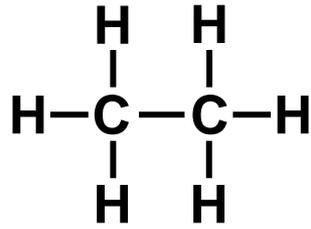
Enlace σ (sigma)

Enlace entre dos átomos en que la distribución de densidad electrónica tiene simetría rotacional alrededor de la línea que une los núcleos de los átomos enlazados.

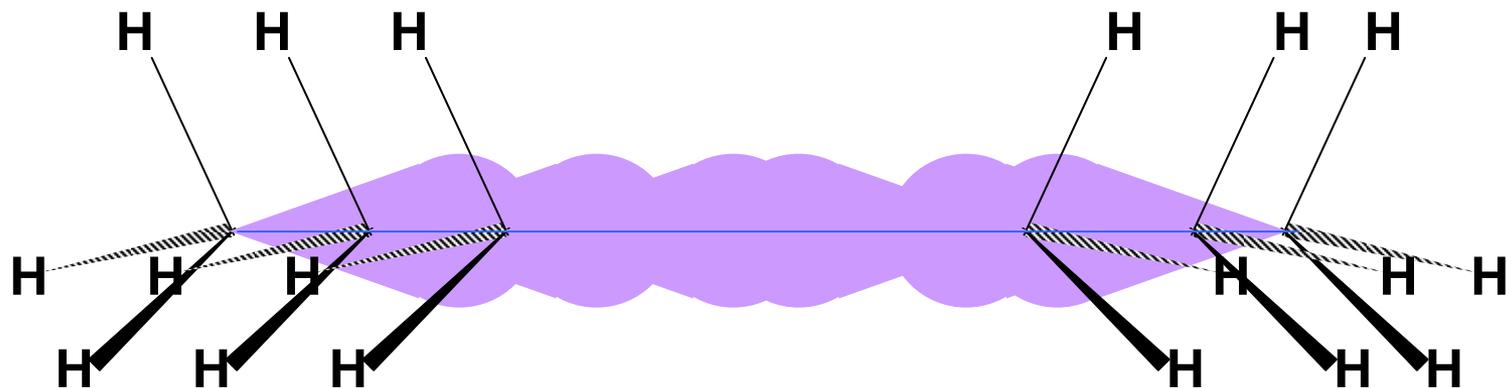
La mayor densidad electrónica se encuentra entre los dos núcleos.



Etano

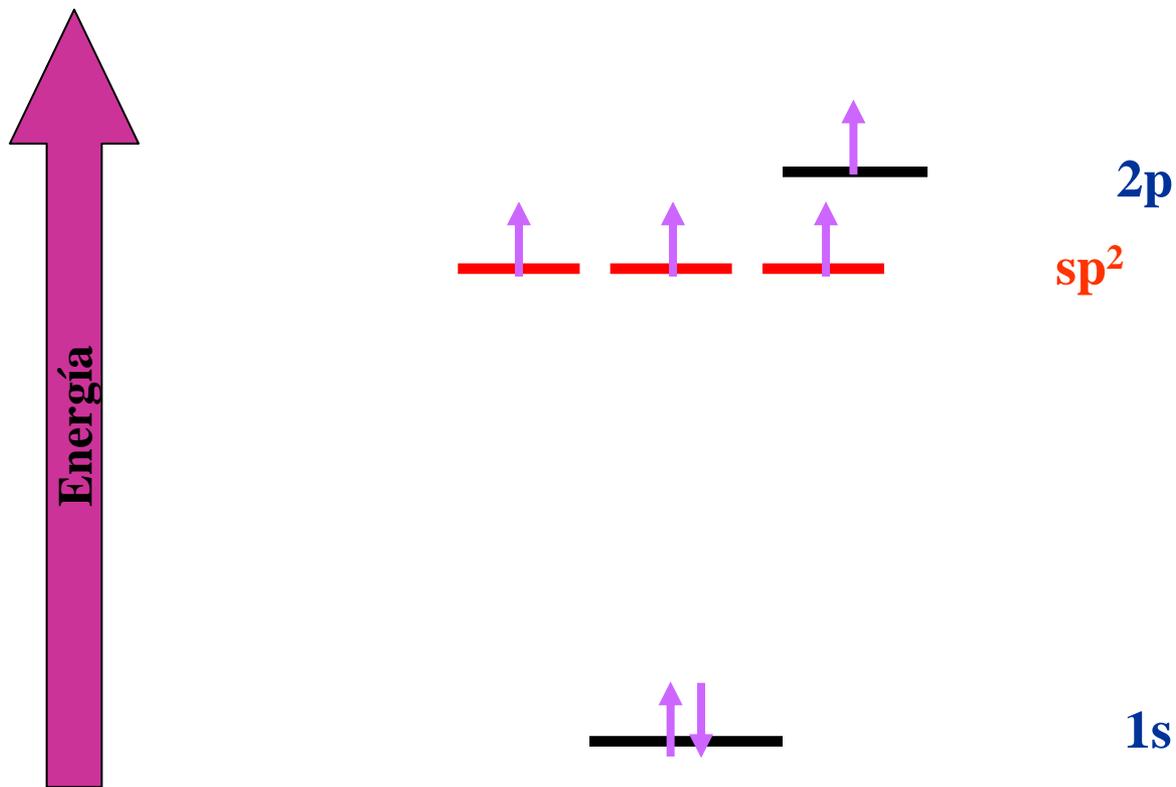


Etano

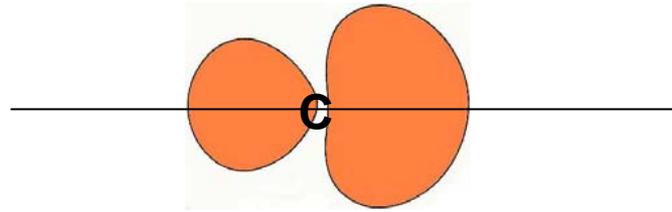


Hibridación de OA

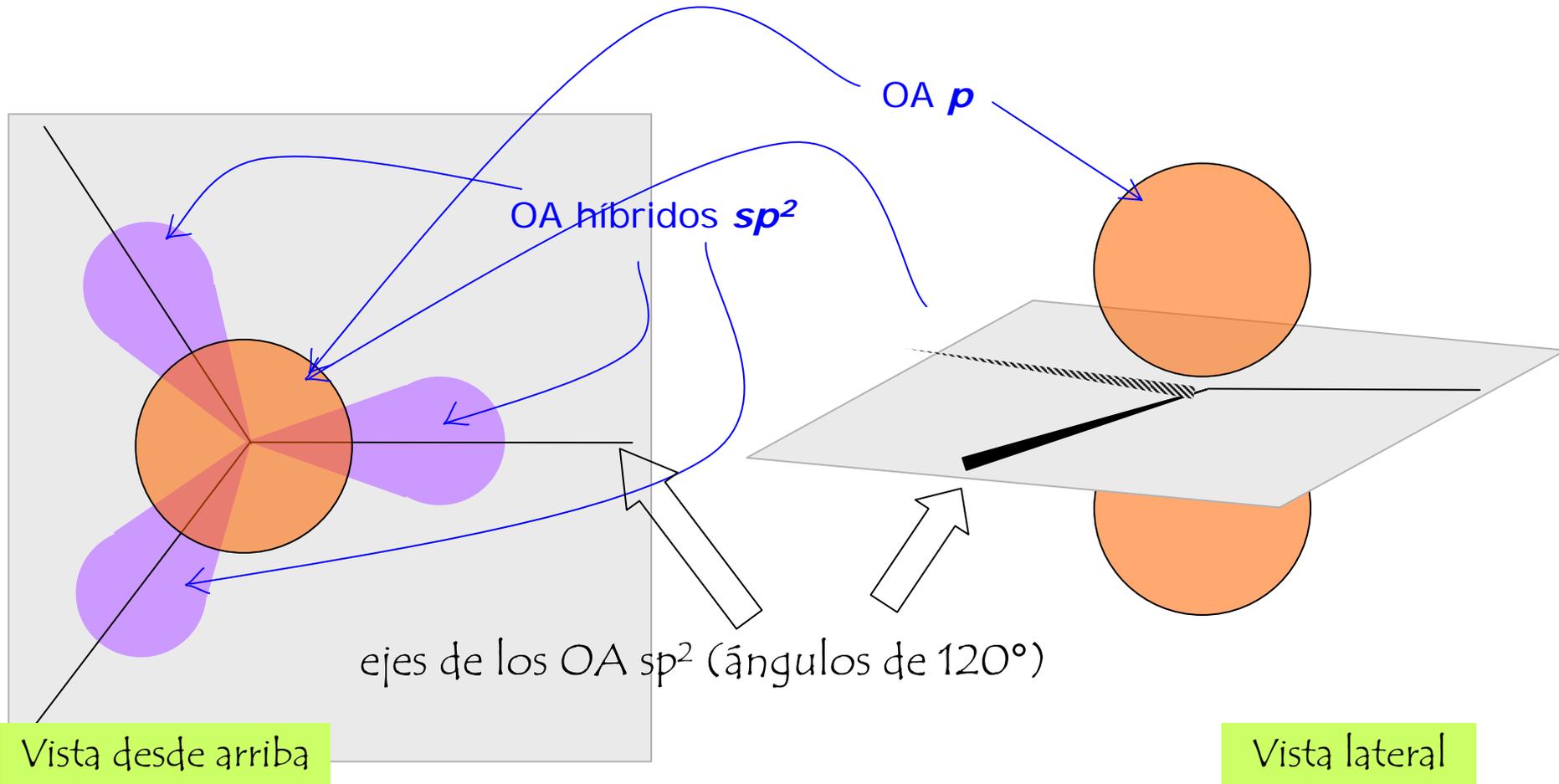
Hibridación sp^2 y la estructura de eteno.



Carbono con hibridación sp^2 : tres orbitales sp^2 y un orbital $2p$.

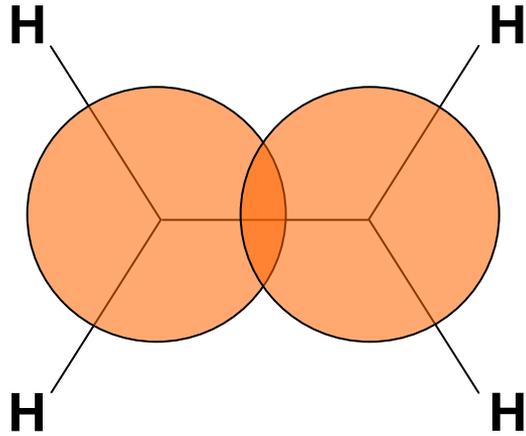
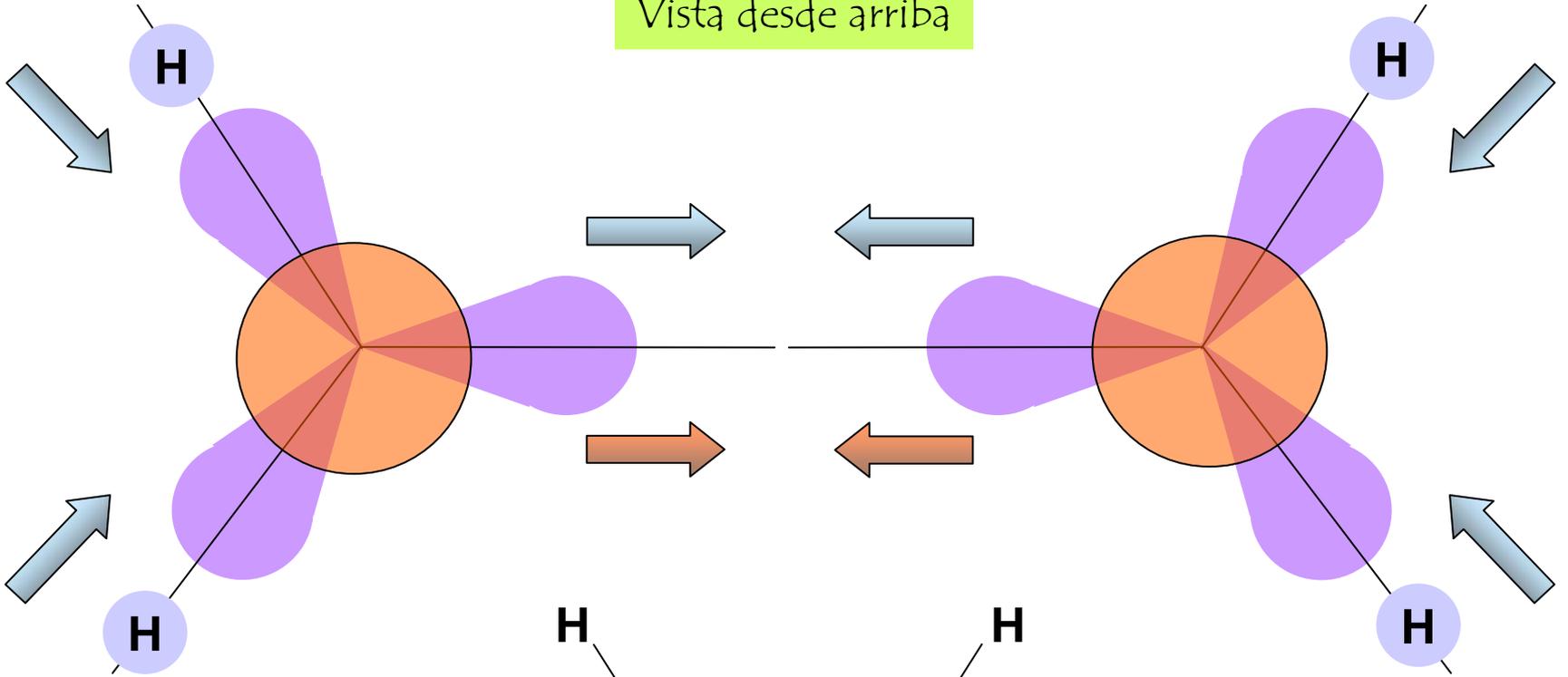


Orbital híbrido sp^2



Eteno

Vista desde arriba

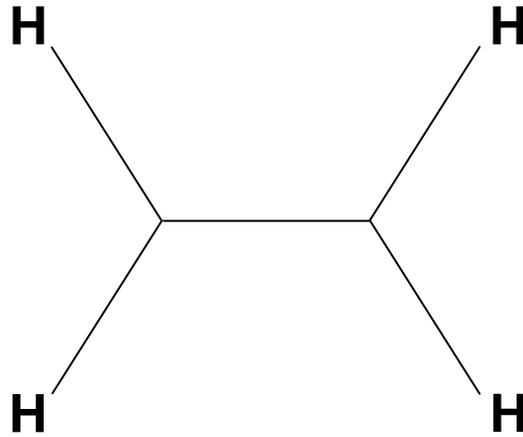
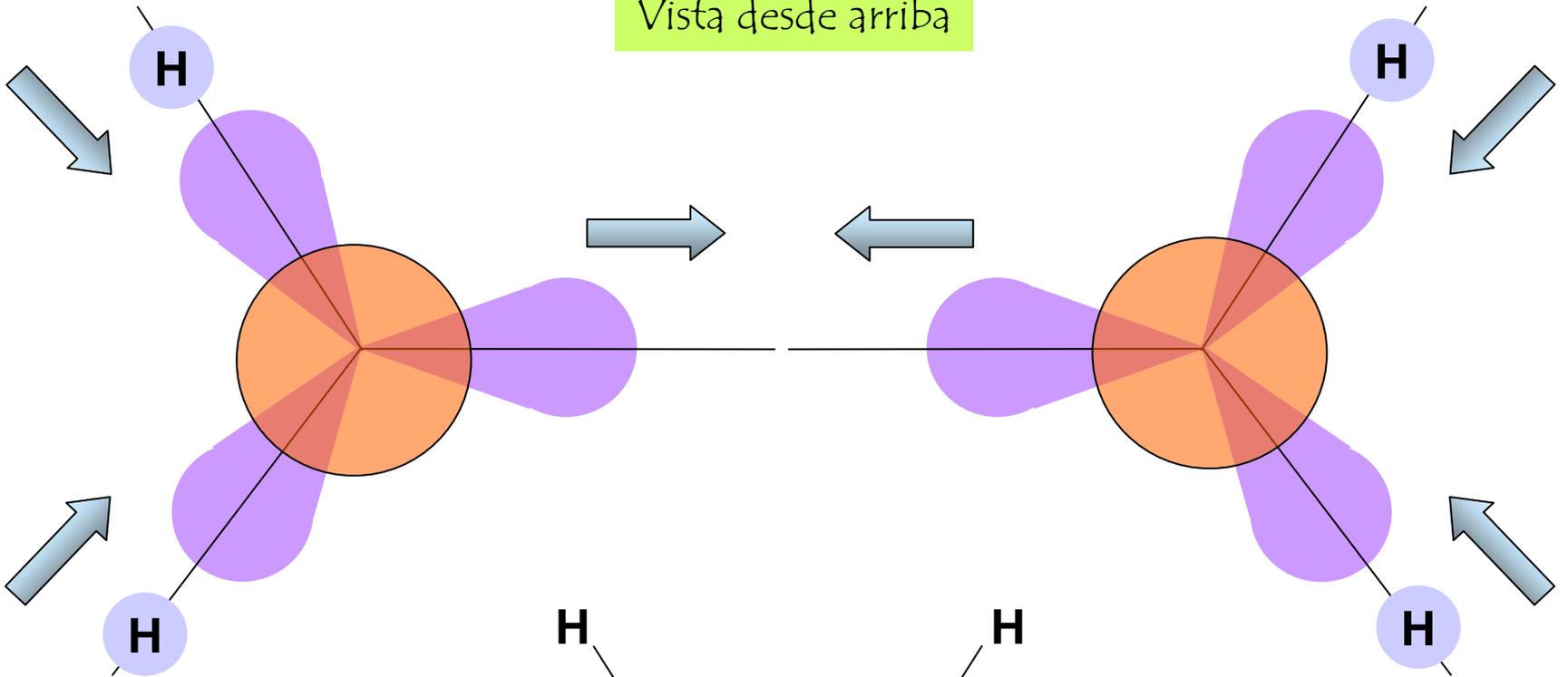


Enlaces σ

Enlace π

Eteno

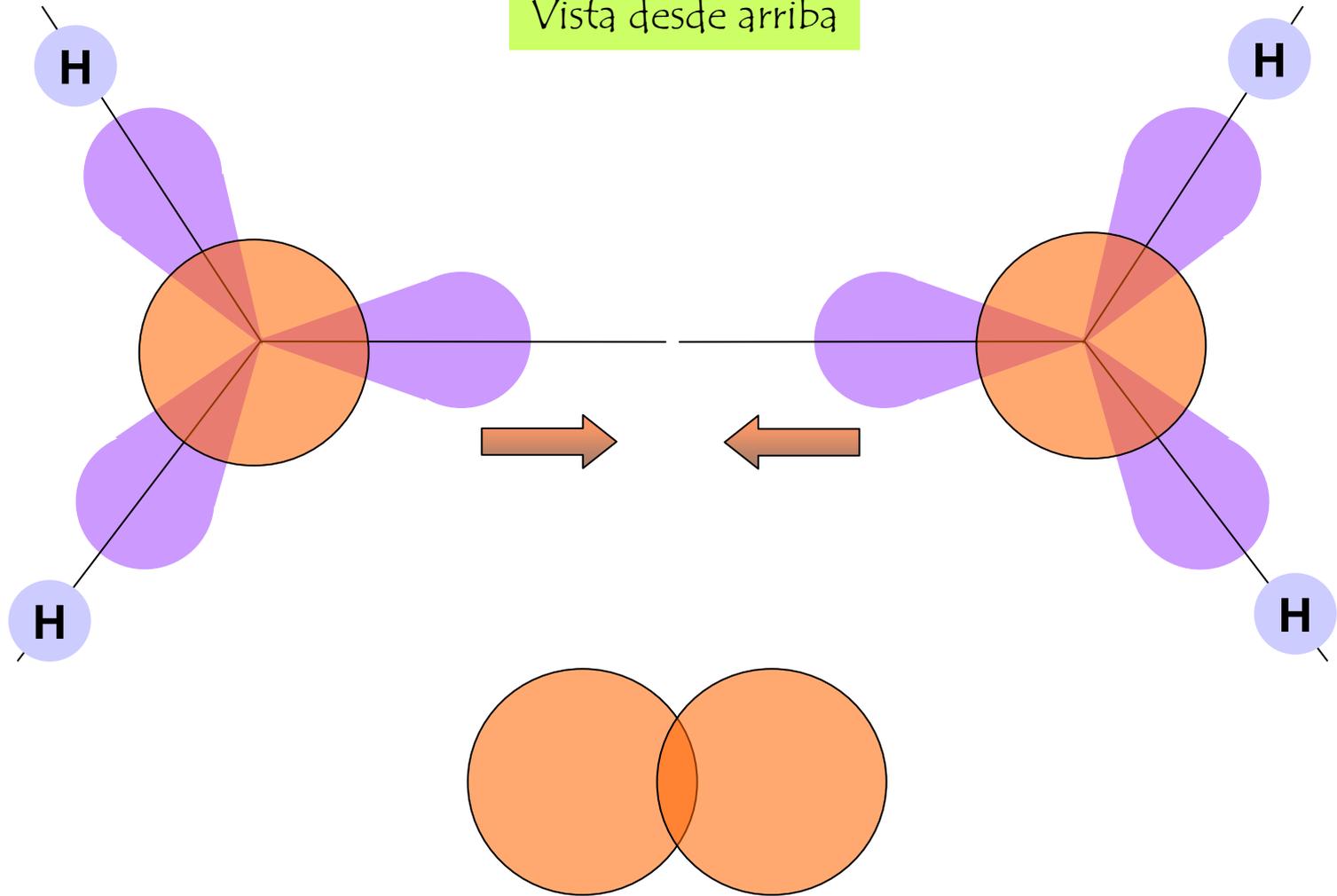
Vista desde arriba



Enlaces σ

Eteno

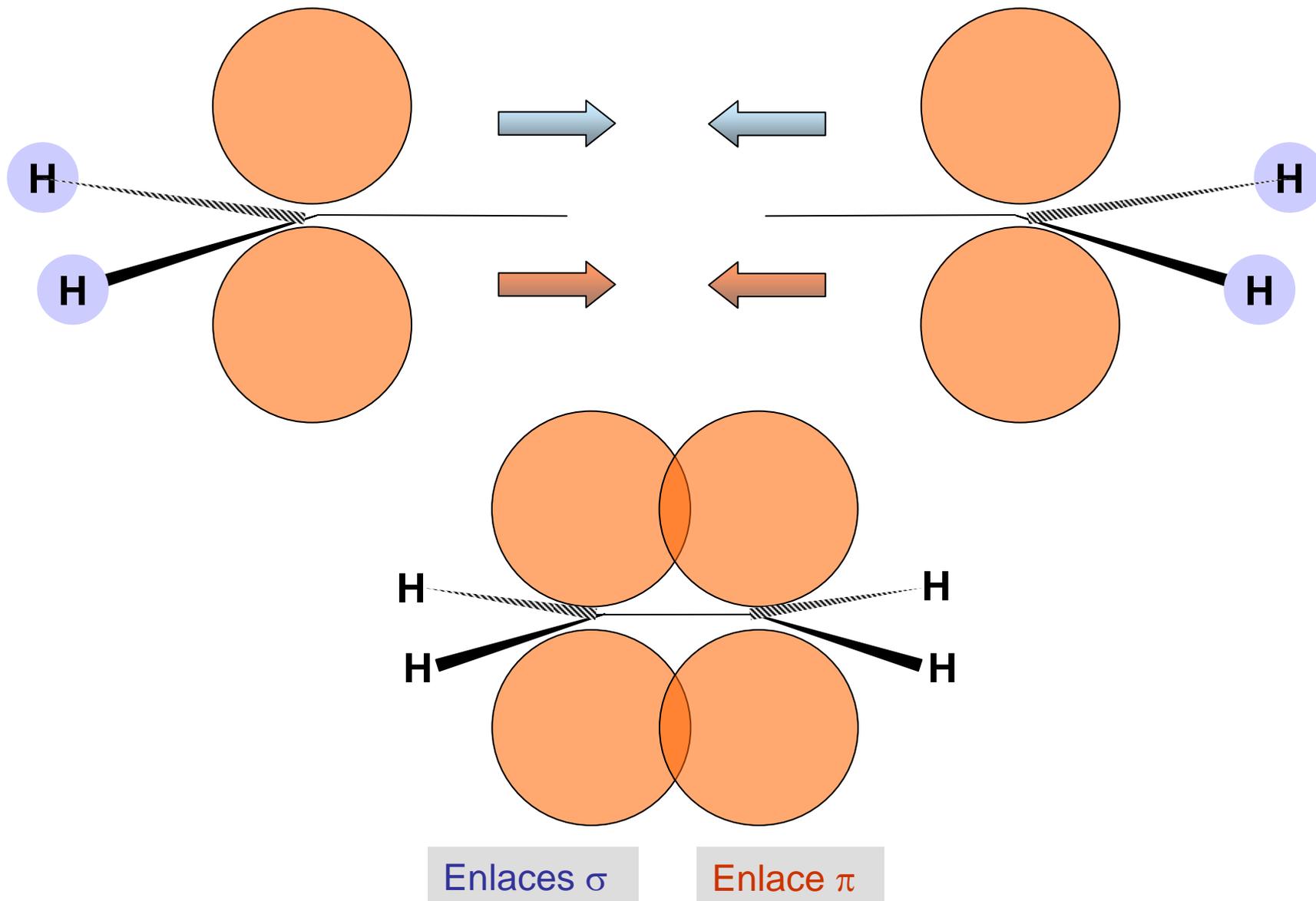
Vista desde arriba



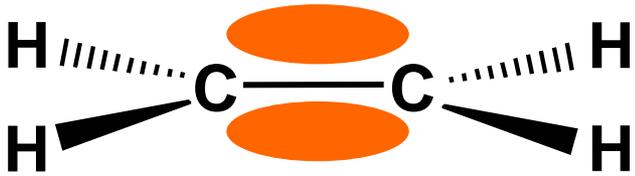
Enlace π

Eteno

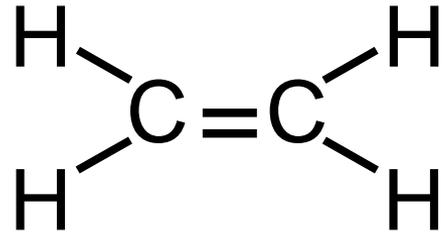
Vista lateral



Eteno



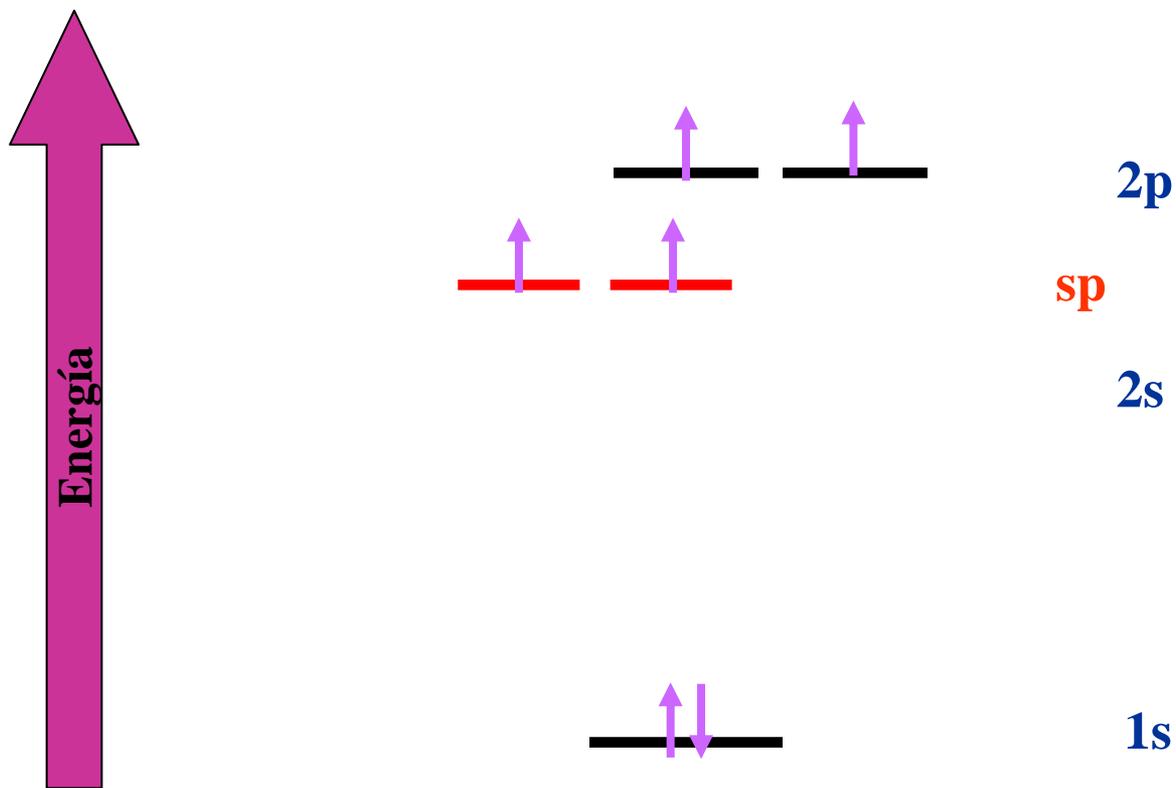
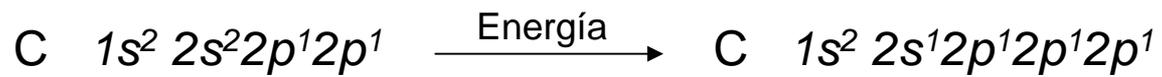
Molécula plana



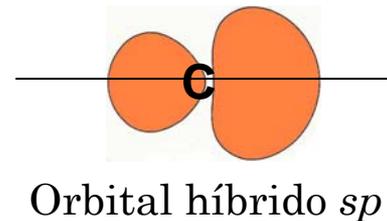
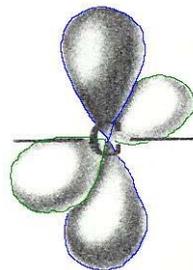
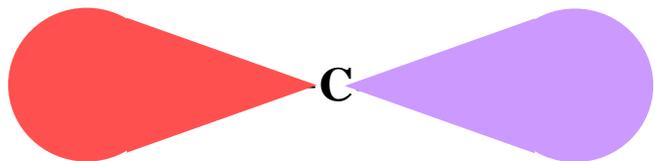
Ângulos de enlace de aproximadamente 120°

Hibridación de OA

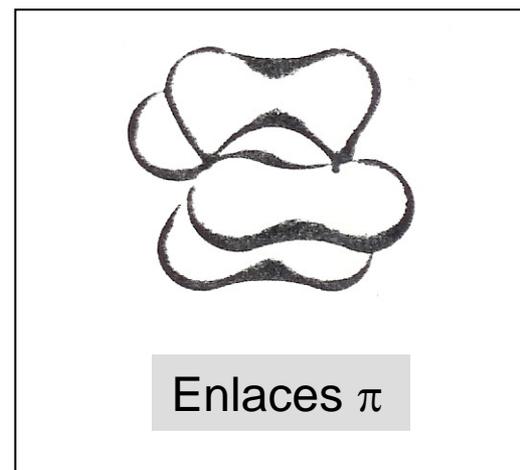
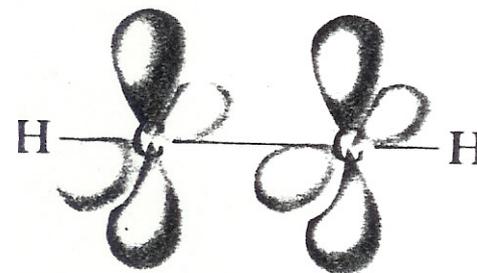
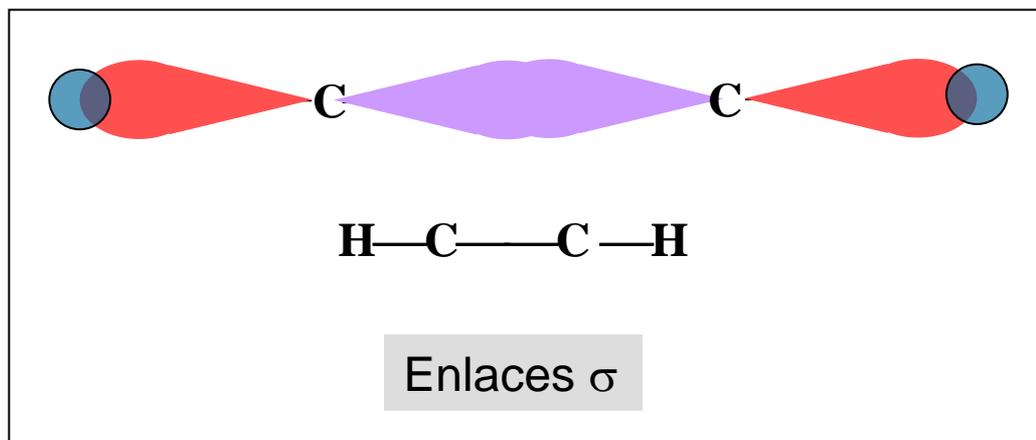
Hibridación sp y la estructura de etino.



Carbono con hibridación sp : **dos orbitales sp y dos orbitales $2p$.**



Etino



En resumen ...

Orbitales híbridos

Comparación de sus energías

